Date Planned : __ / __ / __	Daily Tutorial Sheet-11	Expected Duration : 90 Min
Actual Date of Attempt : __ / __ / __	Numerical Value Type	Exact Duration :_____

1. 80 mL of $\frac{M}{24}K_2Cr_2O_7$ oxidises 22.4 mL H_2O_2 solution. Find volume strength of H_2O_2 solution.

2. Five moles of ferric oxalate are oxidise by how much mole of $KMnO_4$ in acidic medium.

3. Intramolecular redox

$(NH_4)_2Cr_2O_7 \longrightarrow N_2 + Cr_2O_3 + 4H_2O$

Shows a gain and loss of how many electrons

4. How many grams of I_2 are present in a solution which requires 40 mL, of 0.11 N $Na_2S_2O_3$ to react with it.

5. What volume of 2 N $K_2Cr_2O_7$ solution is required to oxidise 0.81 g of H_2S in acidic medium.

$K_2Cr_2O_7 + H_2S + H_2SO_4 \longrightarrow S + K_2SO_4 + Cr_2(SO_4)_3$

6. How many litre of Cl_2 at STP will be liberated by oxidation of NaCl with 10 g of acidified $KMnO_4$.

7. The number of mole of potassium salt i.e. $KHC_2O_4H_2C_2O_4 \cdot 2H_2O$ oxidized by 4 mole of potassium permanganate ion is :

8. 0.2 g of a sample of H_2O_2 required 10 mL of 1 N $KMnO_4$ in a titration in the presence of H_2SO_4. Purity of H_2O_2 is :

9. 5g of a sample of bleaching powder is treated with excess acetic acid and KI solution. The liberated I_2 required 50 mL of $\frac{N}{10}$ hypo. The % of available chlorine in the sample.

10. What is the oxidation no. N in NOCl ?

11. The ratio of oxygen atom having -2 and -1 oxidation no. in $S_2O_8^{2-}$ is :

12. 8 g of Sulphur is burnt to form SO_2 which is oxidized by Cl_2 water. The solution is treated with $BaCl_2$ solution. The number of moles of $BaSO_4$ precipitated is :

13. 25 g of a sample of $FeSO_4$ was dissolved in water containing dil. H_2SO_4 and the volume made upto 1 litre. 25 mL of this solution required 20mL of $\frac{N}{10}KMnO_4$ for complete oxidation. Calculate % of $FeSO_4 \cdot 7H_2O$ in given sample.

14. Calculate the mass of MnO_2 which will be completely oxidises 30 mL of 0.1 M $H_2C_2O_4$ in acidic medium?

15. The action of H_2SO_4 on KI gives I_2 and H_2S. Calculate the volume of 0.2 M H_2SO_4 to produce 3.4 g H_2S.